

Spectrophotometric Investigation of the Interactions between Cationic (C.I. Basic Blue 9) and Anionic (C.I. Acid Blue 25) Dyes in Adsorption onto Extracted Cellulose from *Posidonia oceanica*

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Abstract

Extracted cellulose from *Posidonia oceanica* was used as an adsorbent for removal of a cationic (Basic blue 9, BB) and anionic textile dye (Acid blue 25, AB) from aqueous solution in single dye system. Characterization of the extracted cellulose and extracted cellulose-dye systems were performed using several techniques such as Fourier transform infrared spectroscopy (FTIR), X-ray photoelectron spectroscopy (XPS), zeta potential and Boehm acid-base titration method. Adsorption tests showed that the extracted cellulose presented higher adsorption of BB than AB in single dye system, revealing that electrostatic interactions are responsible, in the first instance, for the dye-adsorbent interaction. In single dye systems, the extracted cellulose presented the maximum adsorption capacities of BB and AB at 0.955 and 0.370 mmol.g⁻¹, respectively.

Adsorption experiments of AB dye on extracted cellulose saturated by BB dye exhibited the release of the latter dye from the sorbent which lead to dye-dye interaction in aqueous solution due to electrostatic attraction between both species. Interaction of BB and AB dyes were investigated using spectrophotometric analysis and results demonstrated the formation of a molecular complex detected at wavelengths 510 and 705 nm when anionic (AB) and cationic (BB) dye were taken in equimolar proportions. The adsorption isotherm of AB, taking into account the dye-dye interaction was investigated and showed that BB dye was released proportionately by AB equilibrium concentration. It was also observed that AB adsorption is widely enhanced when the formation of the molecular complex is disadvantaged.

Keywords: Extracted cellulose; Single dye system; Extracted cellulose saturated by BB dye; Dye-dye interactions; Molecular complex

Introduction

Dyes are complex aromatic substances essential in many industries such as textile, paper, paint, leather and pharmaceutical. Compared with natural dyes [1], synthetic dyes are superior in terms of color availability, easiness to use, quick-setting and ensured by accurate formulas [2]. However, most of the dyes are stable to light, heat and many oxidizing agents, and more difficult to be biodegraded. With the introduction of strict environmental legislation, effluents containing dyes require proper treatment prior to discharge, not only for their high chemical oxygen demand (COD), suspended solids and toxic breakdown products, but also for color, which is not only highly visible and undesirable, but also harm the environment and cause health problems to humans and aquatic animals.

In this context, several methods of decolorization were used such as biological treatment [3], coagulation-flocculation [4] chemical and electrochemical oxidation [5-7]. However, these treatments are generally expensive and their effectiveness is sometimes asked. Sorption onto solid surface, namely adsorption, has proven to be an effective and cheaper process for removing pollutants from textile effluents due to its simplicity and high efficiency. However, the conventional solid retention, usually activated carbons [8] remains for the professionals a relatively expensive material [9]. Recent studies have been conducted to propose alternative sorbents.

Posidonia oceanica (*P.oceanica*), a local biomass abundant on the coasts of Tunisia has shown its effectiveness in removing organic pollutants (textile dyes [10-13]) and inorganic (heavy metals [10,11]). The cellulosic material namely cellulose and hemicellulose, presented the major constituent of the *Posidonia oceanica* (62%) with a relatively high lignin (27%) [14]. During this work, firstly, the extracted cellulose

from the marine plant is used to remove cationic (Basic blue 9, BB) and anionic dye (Acid blue 25, AB) from aqueous solution in single dye system. Secondly, adsorption of an anionic dye (AB) onto extracted cellulose previously saturated by cationic dye (BB) was investigated to understand the adsorption mechanism and the dye-dye interactions on one hand and between the dyes and the adsorbent on the other hand.

Experimental

Adsorbent preparation

Posidonia oceanica's balls were collected from Tunisian coasts. They were washed with tap water to remove sand and other solids, dried in sunlight, crushed into small pieces (45-1000 μm) using a grinder (AM80 Nx2), then washed repeatedly with distilled water and finally dried in an oven at 60°C for 24 h. The cellulose extraction was performed using the follow protocol [15]. The powdered material (20 g) was first treated with 100 mL of a 3 mol.L⁻¹ sodium hydroxide solution at 100°C for 4 h. This alkaline treatment removes both lignin and hemicellulose. The residue was collected by filtration, washed with

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distilled water, and then dried in an oven for 16 h at 50°C. In order to eliminate the pectin as well as some residuals lignin that gives brown colour to the biomass, conventional bleaching treatment was used. In fact, 10 g of the cellulose were agitated with NaClO solution (100 mL; pH 10) at 50°C for 90 min. After filtration, the residue was washed thoroughly with distilled water until the filtrate was neutral, dried in sunlight for two days, and then stored in plastic bottles for further use.

The second adsorbent of this study is extracted cellulose from *Posidonia oceanica* saturated by the basic dye BB. Indeed, the cellulosic material (5 g.L⁻¹) is brought into contact with a solution of dye [BB] = 1 g.L⁻¹ in the aqueous dye solution pH (without addition of acid or base) (pH = 6.2 ± 0.4). The whole is subjected to continuous agitation of 150 rpm (Heidolph Vibramaxe 100) for 24 h. The extracted cellulose is filtered and washed thoroughly with distilled water until the filtrate became clear water. Residual concentrations of BB in all filtrates solutions were determined using the colorimetric method at the maximum wavelength of BB dye (655 nm). Thereafter, the adsorbed amount of BB on extracted cellulose was calculated as 0.586 mmol.g⁻¹.

Characterization of materials

Acidic sites on extracted cellulose were determined by the acid–base titration method proposed by Boehm [16]. Aqueous solutions of NaHCO₃ (0.05 mol.L⁻¹), Na₂CO₃ (0.05 mol.L⁻¹) and NaOH (0.05 mol L⁻¹) were prepared. A volume 25 mL of these solutions was added to 0.5 g of extracted cellulose, shaking for 48 h and then centrifuged. The excess of base was then determined by back titration using NaOH (0.05 mol L⁻¹) and HCl (0.025 mol.L⁻¹). All titrations were carried out at room temperature. The number of acidic sites was determined under the assumptions that NaOH neutralizes carboxylic, lactonic, and phenolic groups; that Na₂CO₃ neutralizes carboxylic and lactonic groups; and that NaHCO₃ neutralizes only carboxylic groups. The test was repeated at least three times.

The Zeta potential of extracted cellulose was measured with Zeta Meter System 3.0. 100 mg of extracted cellulose were added to 100 mL of distilled water. After pH adjustment, mixing for 2 h and centrifugation of the sorbent suspension, the filtrate was placed in the electrophoresis cell for zeta potential measurements. The measurements were repeated at least five times.

Scanning electron microscopy (SEM) analysis was carried out on the extracted cellulose to study its surface texture. The micrograph was recorded on a ZEISS-ULTRA55 SEM device, operating at 10 kV. Prior to the analysis by SEM, the surface was coated with a gold/palladium layer.

Several methods were used to characterize the extracted cellulose, before and after dye adsorption. The FT-IR spectra were recorded in a bio-Rad spectrophotometer. The analysis was applied on different samples to determine the surface functional groups. The acquisition conditions were 64 scans and 4 cm⁻¹ resolution. The wavenumber scanning was in the range of 500-4000 cm⁻¹.

The X-ray photoelectron spectroscopy (XPS) experiments were performed with a XR3E2 apparatus (Vacuum Generators, UK) equipped with a monochromated MgKα X-ray source (1253.6 eV) and operating at 15 kV under a current of 20 mA. Samples were placed in an ultrahigh vacuum chamber (10⁻⁸ mbar) with electron collection by a hemispherical analyzer at an angle of 90°. The signal decomposition was realised using Spectrum NT, and the CAH signal was used as a reference peak at 285.0 eV. XPS analysis was performed in extracted cellulose before and after dye adsorption.

Dyes and analysis

AB (molecular formula = C₂₀H₁₄N₂O₅S, λ_{max} = 602 nm) and BB (molecular formula = C₁₆H₁₈N₃SCl, λ_{max} = 655 nm) were supplied by Reactifs–RAL and used without further purification. The characteristics and structure of these dyes is presented in Figure 1. The test solutions were prepared by diluting stock solutions to the desired concentrations. The UV–vis spectra of aqueous dye solutions were measured by a double beam UV/vis spectrophotometer (Cecil instrument, Model CE2021-2000).

Batch sorption studies

A fixed amount of extracted cellulose (0.25 g) was added to a set of bottles (Teflon) containing 50 mL of known concentration of AB and BB dyes in single solution at pH around 6.2 (without addition of acid or base). The bottles were agitated at 150 rpm (Heidolph Vibramaxe 100 shaker) and at temperature of 21°C until equilibrium was reached. Then, samples were centrifuged at 2000 rpm for 2 min, and the supernatant was analyzed for the residual AB and BB concentration (by measuring the absorbance at 602 and 665 nm for AB and BB, respectively) using a double beam UV/VIS spectrophotometer (Cecil instrument, Model CE2021-2000).

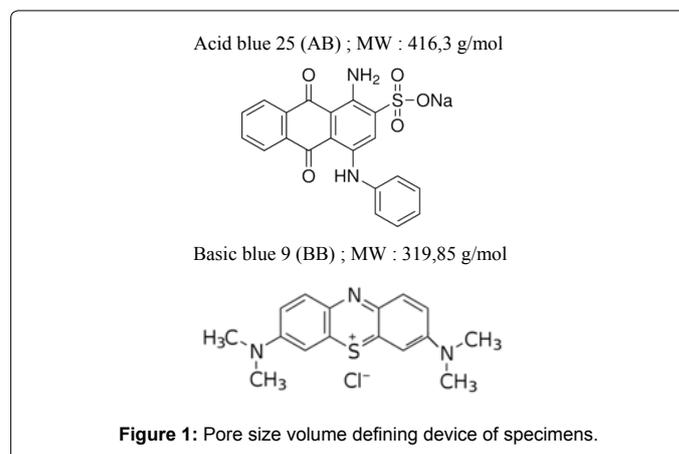
On the other hand, bottles' series of AB solutions (50 mL) at different initial concentrations (0-1 g.L⁻¹) were placed in contact with 0.25 g of extracted cellulose saturated by the basic dye BB. The flasks were agitated at 150 rpm (at temperature of 21°C) until equilibrium was reached and centrifuged at 2000 rpm for 2 min. The residual dye solutions after adsorption were scanned over the whole visible range. In this work, the classical equations Freundlich and Langmuir were used to model experimental curves and the equilibrium models were fitted employing the non-linear fitting method, using the non-linear fitting facilities of the software Microcal Origin 6.0.

Results and Discussion

Characterization of extracted cellulose

Figure 2 shows the SEM micrograph of the extracted cellulose sample. The extracted cellulose has considerable numbers of heterogeneous layers and pores which may be very interesting for dye adsorption.

Table 1 summarizes the results of Boehm titration of extracted cellulose and shows that most of acidic functional groups are carboxylic, followed by phenolic and then lactonic.



The wide survey scan spectra are shown in Figure 3. The spectrum of extracted cellulose shows two strong symbol peaks for C1s and O1s at 285 and 531 eV respectively. Three quite weak ones for N 1s, Ca 2p3/2 and Cl 2p3/2 can be observed at, 395.7, 345.7 and 200 eV, respectively [17]. Those observations suggest the presence of abundant C, O and minor N, Ca and Cl on the surface of the extracted cellulose. The elemental surface composition of extracted cellulose obtained by XPS is given in Table 2, which shows that the surface of the extracted fiber did not consist of pure cellulose. The O/C ratio was much lower than in pure cellulose (0.83 in theory). Some nitrogen, calcium and chlorine were observed, which should not be present in pure cellulose. The calcium atom has been often detected on natural celluloses [18] and the chlorine atom could be present from bleaching operation. Those results confirm that the extracted cellulose yet contains lignin and hemicelluloses.

The FT-IR spectrum of extracted cellulose shows a strong absorption band at 3430 cm^{-1} due to stretching of O-H groups and

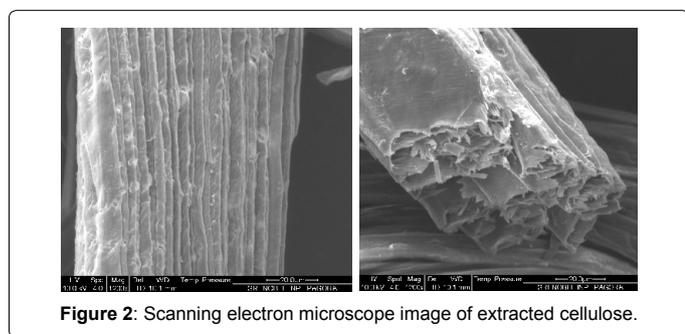


Figure 2: Scanning electron microscope image of extracted cellulose.

pH _{PZC}	Acidic groups (mmol.g ⁻¹)			
	Carboxylic	Lactonic	phenolic	total
3.5	0.875	0.015	0.275	1.165

Table 1: pH_{PZC} value and Boehm titration results of extracted cellulose.

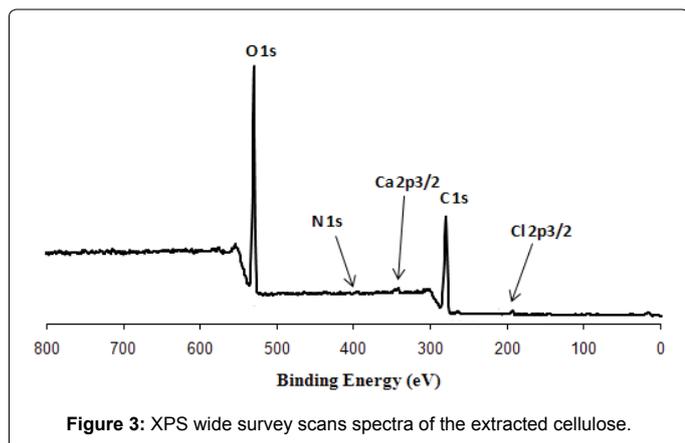


Figure 3: XPS wide survey scans spectra of the extracted cellulose.

	Surface composition (at percentage)						O/C	S/C
	O	C	N	Ca	Cl	S		
Extracted cellulose	35.62	61.78	0.68	0.41	0.67	-	0.57	-
Extracted cellulose after AB adsorption	36.17	62.90	0.91	0.09	0.13	0.04	-	6.3.10 ⁻⁴
Extracted cellulose after BB adsorption	28.20	67.93	2.33	0.17	0.71	0.08	0.42	-

Table 2: The elemental surface composition from XPS wide survey spectra.

that one at 2891 cm^{-1} to the C-H stretching. The band at 1651 cm^{-1} corresponds to the bending mode of the adsorbed water [19] but can also be assigned to the absorption of carboxylic groups.

Dye adsorption in single-dye solutions

Effect of the solution pH on dye adsorption: The influence of equilibrium pH on the BB and AB sorption was investigated between 3 and 11 and the results are shown in Figure 4. Dyes sorption onto extracted cellulose appeared to be strongly pH-dependent. Increasing the pH value from 3 to 4 resulted in a considerable increase in BB sorption and higher decrease in AB one, whereas further increasing the pH value from 4 to 10 yielded a slight change in sorption, and this for both dyes. This tendency correlated well with the variation of surface charge of the extracted cellulose (the pH of point zero charge pH_{ZPC} measured by electrophoresis is 3.5). Consequently, for pH values lower than pH_{ZPC} the sorbent exhibits a positive surface charge. The higher amount of anionic dye (AB) sorbed at pH values lower than 3.5 can be explained by the electrostatic interactions between the positively charged surface of the sorbents and of the negatively charged AB dye. For pH higher than 3.5 the extracted cellulose surface is negatively charged. This further indicates that cationic dye (BB) sorption was also driven primarily via electrostatic interaction.

Equilibrium sorption isotherms: The sorption studies of AB and BB dyes in single dye solutions onto extracted cellulose were carried out in a batch reactor by varying the initial concentrations of the studied dyes at an initial pH of 6.2 ± 0.4. Nonlinear regression was used to fit the dye sorption isotherm to the Freundlich and the Langmuir (Table 3). Based on the R² values the best isotherm model fitted for extracted cellulose sorbents is the Langmuir model (Figure 5). Adsorption capacities of BB and AB on the extracted cellulose were determined to be 0.955 and 0.370 mmol/g, respectively. The successful application of this model to the present data supports that all adsorption sites are energetically and sterically independent of the adsorbed amount [20]. The solid is assumed to have a limited adsorption capacity, which defines sorbent's total capacity.

Dye-cellulose interaction characterization

FTIR study: In order to illustrate interaction between, firstly BB and extracted cellulose and secondly AB and extracted cellulose, to suggest responsible sites for adsorption, analysis FTIR of this sorbent after adsorption were investigated. Figure 6 shows the FTIR spectra of BB, AB and extracted cellulose before and after dye adsorption. Figure 6a shows the extracted cellulose spectra. For BB (Figure 6b), stretching vibration of tertiary amine groups is observed at 3422 cm^{-1} . Bands corresponding to the C-H stretching vibration are at 2919 and 2849 cm^{-1} . The stretching vibration of nitrile groups is at 2700 cm^{-1} . Vibrations of the aromatic ring are found at 1597 and 1538 cm^{-1} . The C-N stretching vibration is at 1337 cm^{-1} . The C-H 'in plane' vibrations are found between 1253 and 1037 cm^{-1} and 'out of plane' vibration is at 945 cm^{-1} [21].

Similarly, for AB Figure 6c reveals the presence of several peaks of AB's characteristic functional groups. The stretching vibrations observed at 3410 and 1185 cm^{-1} mark the presence of amine groups, band corresponding to the CH stretching vibration is at 2920 cm^{-1} , the functional groups of the aromatic ring such as C=C and C=O are located at around 1560 cm^{-1} , the CN stretching vibration is at 1360 cm^{-1} , CH vibration is at 1262 cm^{-1} and the band at 1017 cm^{-1} can be assigned to the S = O bond according to Auta and Hameed [22].

It is noted that the shoulder observed between 1100 and 1000

cm^{-1} in cellulose spectrum, which corresponds to the deformation vibrations in the plane of hydroxyl groups of alcohols, is modified after incorporation of the dyes. The peaks that appear in (d) and (e) to 1341 cm^{-1} correspond to the CN stretching vibration at each dye (Figures 6b and 6c). In addition, after adsorption of AB (Figure 6d) the band at 1020 cm^{-1} , which corresponds in the case of raw material (a) to the CO stretching vibration has changed appearance at the cellulose adsorbed AB (d) marking the presence of S=O bond at 1017 cm^{-1} . Two other peaks appear at the shoulder at 1128 and 1080 cm^{-1} which probably correspond to the involvement of the carbon group C2 and C6 of the representative unit of the cellulose after dyes adsorption.

Besides peaks' shift positions, the intensity and sharpness of the bands observed for the raw material were also affected after dyes adsorption. The broadband at 3420 cm^{-1} of extracted cellulose (Figure 6a) becomes slightly narrow after dyes adsorption (Figures 6d and 6e), indicating the possibility of the involvement of hydroxyl groups of the extracted cellulose in the adsorption process.

Thus, it can be concluded that the OH group of the material interacts with the nitrogen atom on the BB or AB with hydrogen

bonding [21]. Since the N atom in BB is linked to the aromatic rings or to the terminal methyl groups, or in AB is linked to the primary or secondary amine, the vibrational frequencies of the aromatic ring and other groups will be affected when nitrogen is involved in a hydrogen bonding interaction.

X-ray photoelectron analysis (XPS): After dyes adsorption atomic composition of the surface varied considerably. After BB sorption, increasing of C, N and Cl amounts and decreasing O and Ca one (Table 2), and the appearance of S 2p_{3/2} at 164 eV binding energy was observed. The increase in C, S, N and Cl atomic composition come from the presence of BB species at the surface. While, increasing of C, O and N amounts and decreasing Cl and Ca one (Table 2), and the appearance of S was observed after AB adsorption.

It can be seen from BB molecule formula, that this latter present no oxygen atom, thus, it was interesting to compare O/C ratio of extracted cellulose and the resulting material after adsorption (Table 2). Reciprocally, S/C ratio of AB and extracted cellulose adsorbed the anionic dye was compared. From these values, it is easy to calculate the adsorption capacity of each dye and it could be measured as 2.05 and 0.34 mmol.g^{-1} for BB and AB, respectively. The adsorption capacity (from adsorption isotherm study) of samples analyzed with XPS was measured as 0.955 and $0.0465 \text{ mmol.g}^{-1}$ for BB and AB, respectively, which is lower than the calculated one from XPS results. The AB and BB species might exit preferentially at the surface of the fiber rather than in its porosity.

AB adsorption on extracted cellulose loaded with MB dye

Highlighting dye-dye interactions: The visible absorption spectra of AB and BB were used for the analysis of dyes in single dye solutions. The single solutions of AB and BB containing 20 and 10 mg L^{-1} , respectively of dyes were prepared and the visible spectra of these solutions were recorded between 400 and 800 nm (Figure 7). As it can be seen from Figure 7, maximum wavelength (λ_{max}) of AB and BB were determined as 602 nm and 665 nm, respectively. The calibration curves were prepared at λ_{max} of each dye. On the other hand, this figure shows the addition of absorbance spectra of single solutions of AB and BB (superposition spectra) which is not in agreement with the absorbance spectra of binary solution made by 20 mg L^{-1} of AB and 10 mg L^{-1} of BB. This shows that the Beer-Lambert rule is not applicable due to possible interaction of dyes. So that, derivative spectrophotometry method [23], which is based on the analysis of derivative spectra, cannot therefore be applied in this case to determine the composition of a binary mixture of dyes. A series of binary solutions was prepared to cover the wide

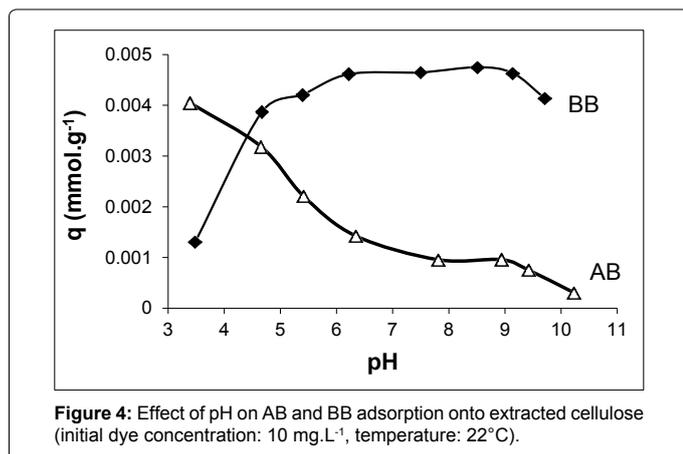


Figure 4: Effect of pH on AB and BB adsorption onto extracted cellulose (initial dye concentration: 10 mg.L^{-1} , temperature: 22°C).

Dye	Langmuir			Freundlich		
	q_m (mmol.g^{-1})	K_L (L mmol^{-1})	R^2	K_F ($\text{mmol}^{1-n} \text{L}^{1/n} \text{g}^{-1}$)	n	R^2
BB	0.955	11.64	0.99	2.22	1.56	0.98
AB	0.370	0.95	0.95	0.17	1.49	0.91

Table 3. Isotherm parameters for AB and BB sorption on extracted cellulose.

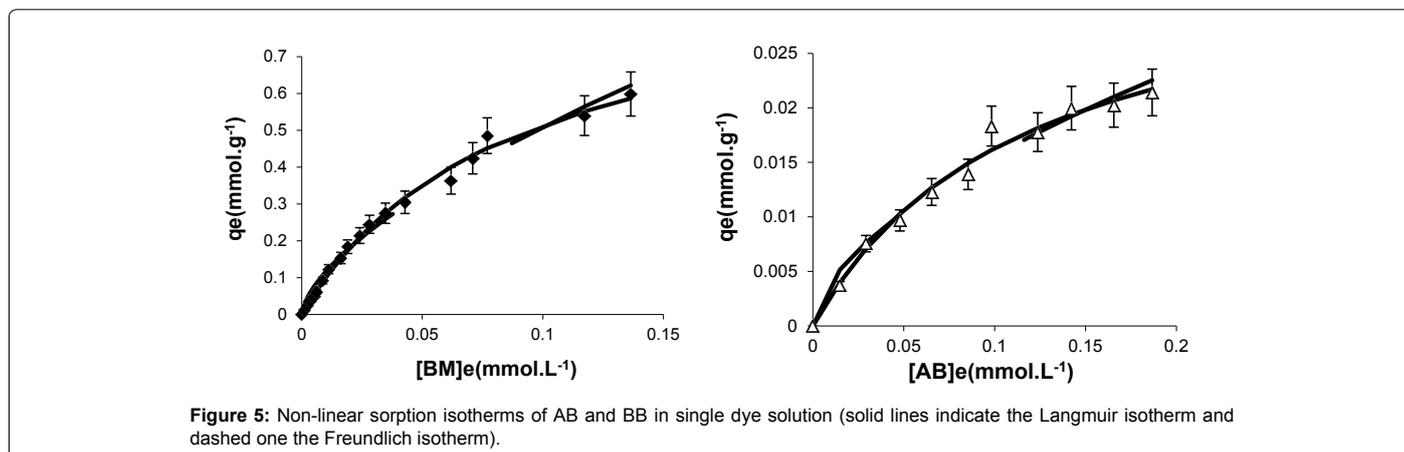


Figure 5: Non-linear sorption isotherms of AB and BB in single dye solution (solid lines indicate the Langmuir isotherm and dashed one the Freundlich isotherm).

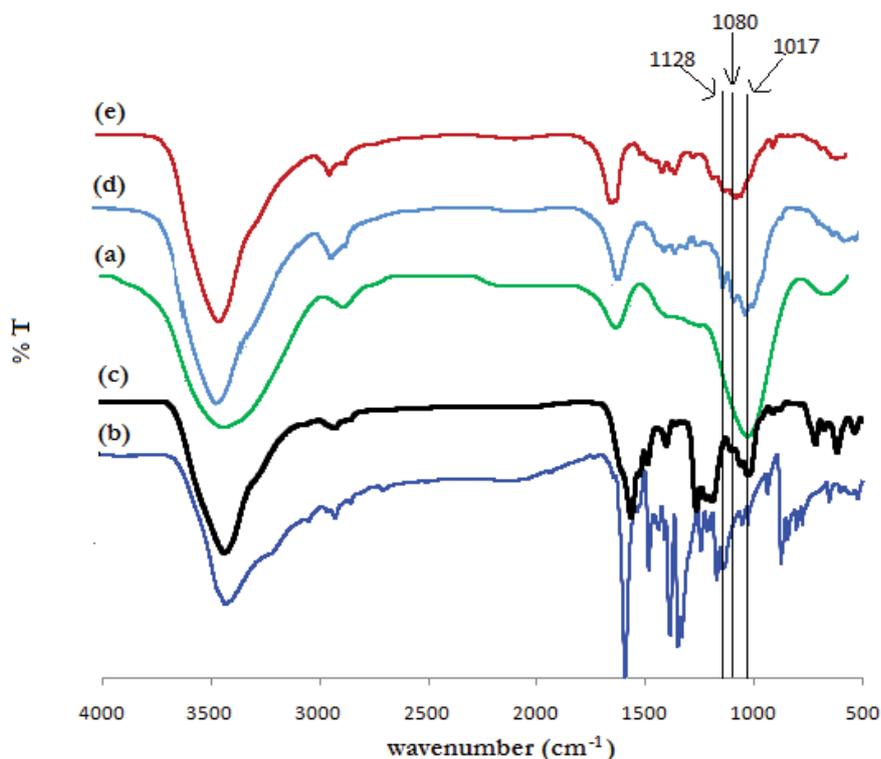


Figure 6: FT-IR spectra of (a) extracted cellulose, (b) BB, (c) AB, (d) extracted cellulose after AB adsorption and (e) extracted cellulose after BB adsorption.

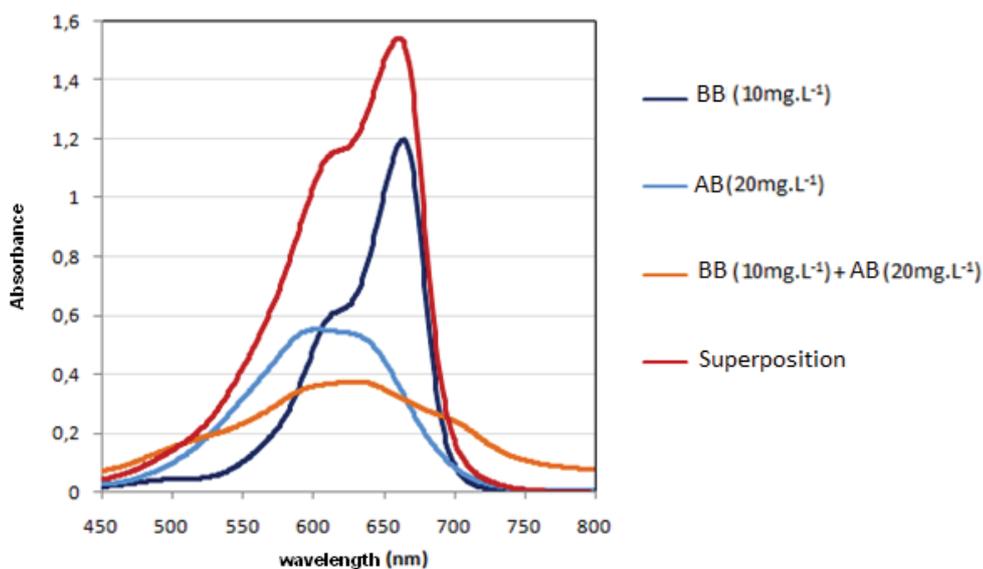


Figure 7: Visible absorption spectra of AB and BB in single and binary solutions.

range of concentrations which will allow, for identification of spectra as fingerprints to determine the composition of the mixtures in residual solutions after dye adsorption. Figures 8-10 show the spectra obtained for a wide range of binary composition mixtures.

The data spectra allow for some findings:

- If $[AB] < [BB]$ the mixture spectrum has a similar shape to that of BB spectrum (Figure 8). It is thus found that for a constant concentration of BM (eg. $[BB] = 0.626 \text{ mmol.L}^{-1}$), the absorbance at 665 nm decreases as AB concentration increases (provided that $[AB]$ remains below $[BB]$) and the shape of the spectrum is closer to that of BM spectrum.

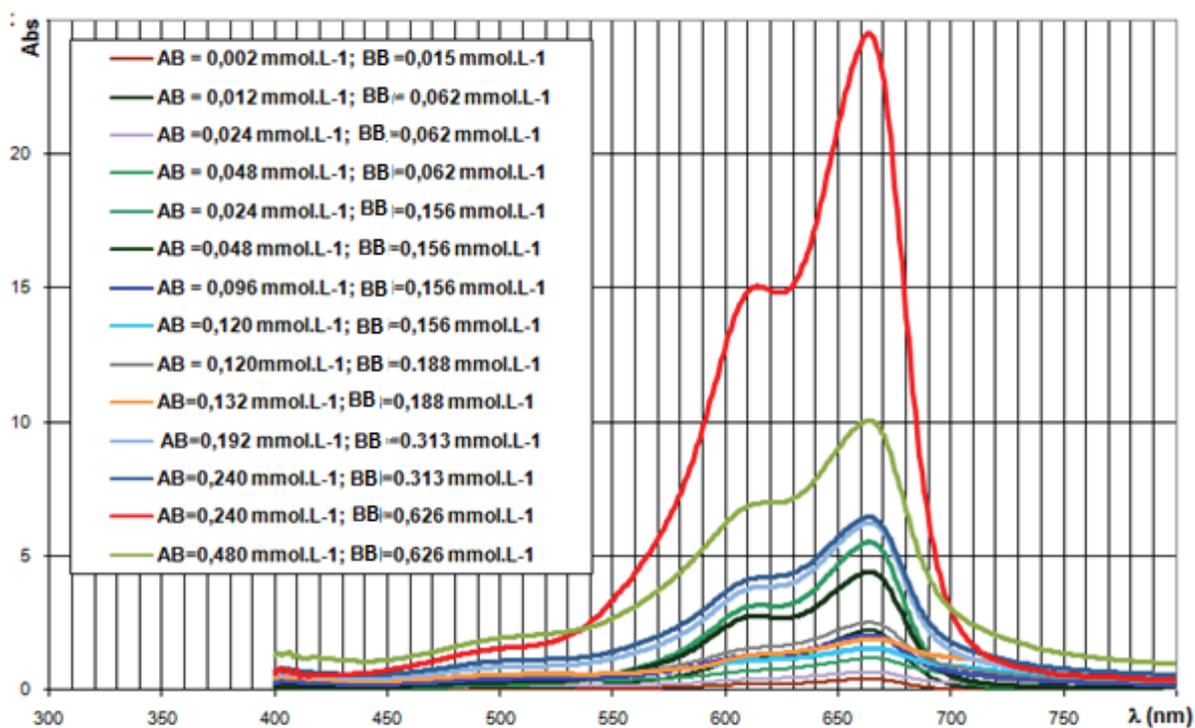


Figure 8: Spectral scanning binary mixtures for which $[AB] < [BB]$.

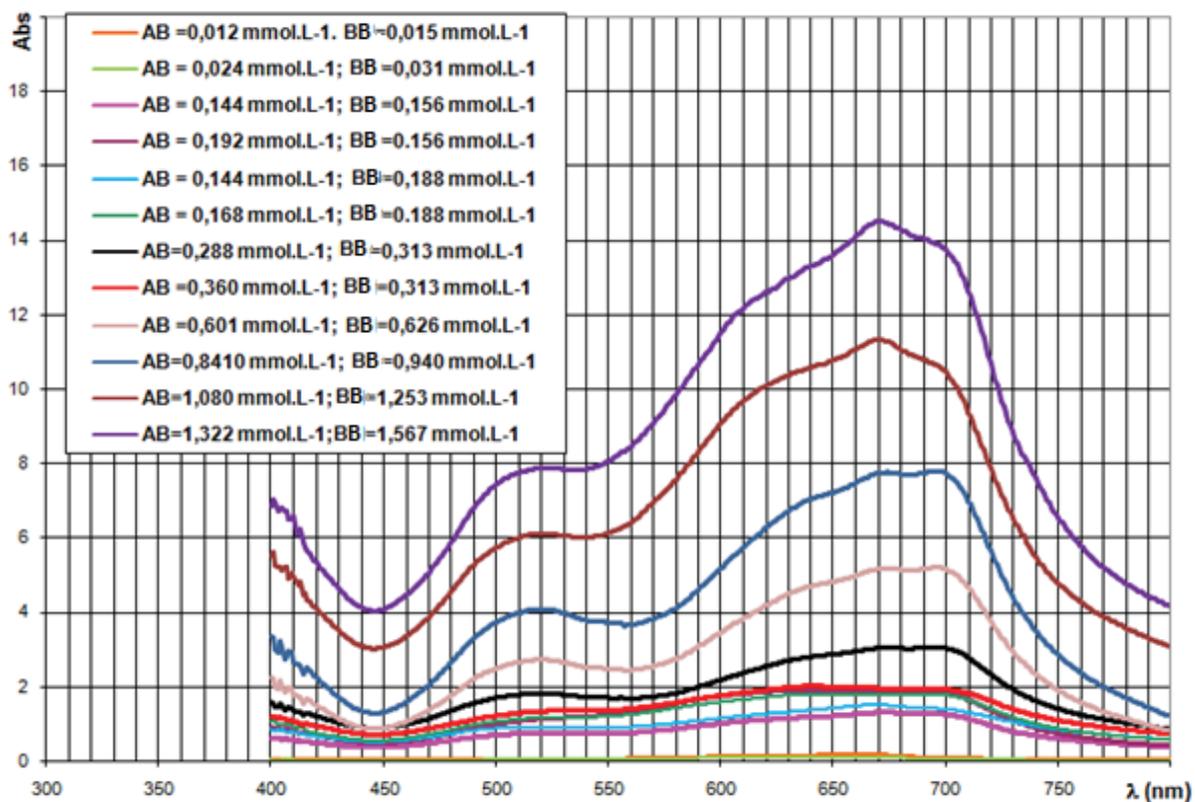


Figure 9: Spectral scanning binary mixtures for which $[AB] = [BB]$.

- If $[AB] > [BB]$ the mixture spectrum has rather the shape of AB spectrum (Figure 10).
- If $[AB] \approx [BB]$ the mixture spectrum has a third different shape from that of the two dyes, with the appearance of shoulders at 705 and 510 nm in addition to the maximum peak wavelength of the two dyes (Figure 9). It should be noted that in this case the absorbance of the shoulder at 705 nm is proportional to both concentration of BB and AB. This suggests an intermediate form takes place when both dyes are present in equimolar amounts. This form, which may be a molecular complex, is detected at 705 and at 510 nm.

Adsorption study: In this study, adsorption of the anionic dye (AB) was carried out after prior adsorption of the cationic dye (BB). The aim is to change the charge state of the adsorbent surface by the binding of the cationic dye and thus enable better sorption of anionic dye. The equilibrium adsorption of AB on the raw extracted cellulose and previously saturated one by BB dye was presented in Figure 11. On the practical side, the amounts of adsorbed AB dye in the adsorption isotherm study onto extracted cellulose saturated by BB dye (O) were determined following a direct reading of the absorbance at 602 nm (λ_{max} of AB) of residual solutions without taking account of the BB interference. Compared to adsorption isotherm of AB obtained on the

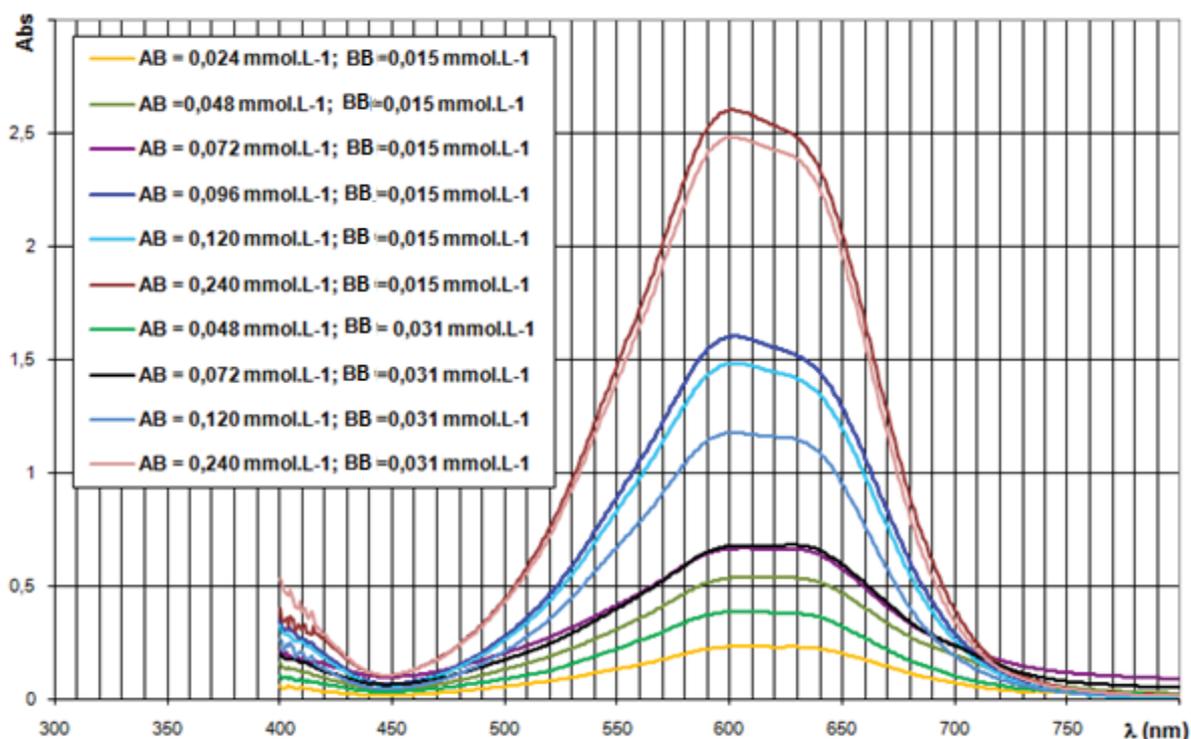


Figure 10: Spectral scanning binary mixtures for which $[AB] > [BB]$.

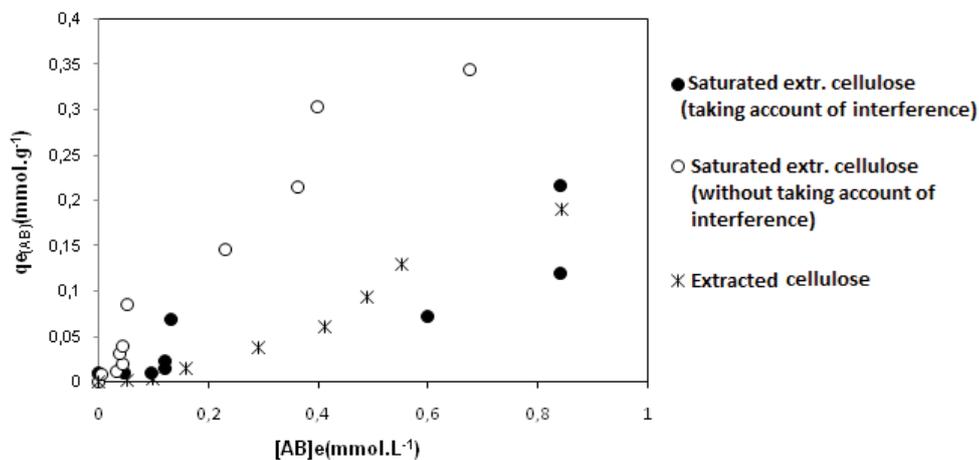
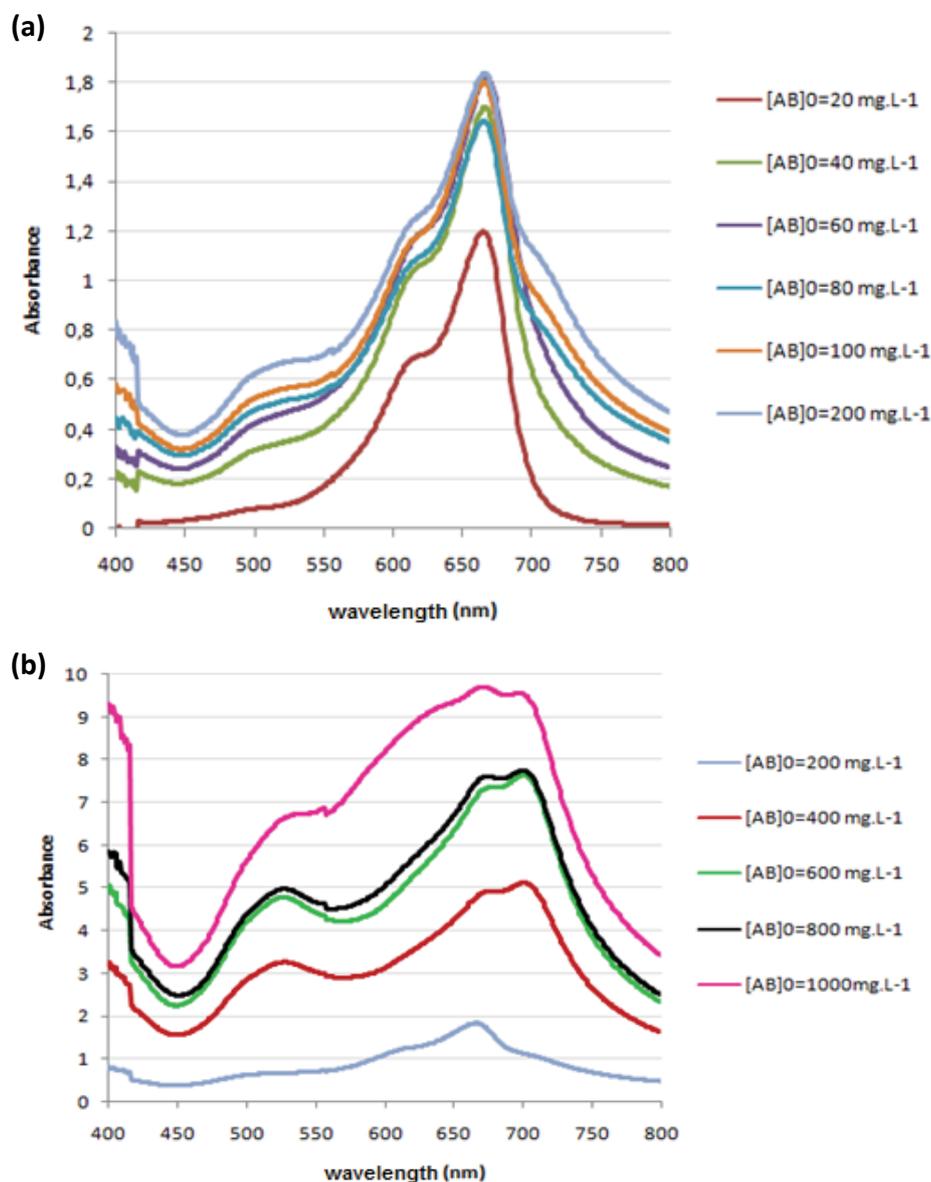


Figure 11: Comparison of adsorption isotherms of AB.

raw extracted cellulose (\mathcal{K}), it is noted that the quantities of adsorbed AB onto extracted cellulose previously saturated with BB dye increased significantly. Thus, we can easily conclude that cationic dye (BB) helped to modify the surface state of the extracted cellulose by increasing the number of active sites for the anionic dye and thereby improving the adsorption. However, Figures 8-10 have demonstrated the effect of the presence of a dye on the spectrophotometric analysis of the other dye. Besides the probable BB release in the adsorption case of AB affects the analysis of the latter and will be the cause of wrong results interpretation. So that, the residual dyes solutions after adsorption were scanned over the entire visible range and the results are shown through Figure 12. The recorded spectra exhibit an absorption maximum at 665 nm and a shoulder at 610 nm (Figure 12a). The shape of these spectra corresponds to BB one, which proves the presence of the latter which is released during AB adsorption.

Moreover, Figure 12b shows that greater the initial AB concentration increases (which corresponds to the increase of the residual concentration) more the shape of the spectra is transformed with the appearance of a shoulder at 520 nm which appears on the dye mixture spectra (Figure 7). Thus, the identification of the solutions composition after adsorption was carried out by superposing the spectra obtained after adsorption (Figure 12) to those presented through Figures 8-10 and which correspond to the solutions at known composition. Consequently, it will be possible to determine AB adsorbed capacity and draw again the adsorption isotherm of the AB dye by taking into account the interfering presence of BB dye (Figure 11(●)).

Thus, Figure 11 highlights the error made by a direct reading of the AB absorbance neglecting the presence of BB and shows that adsorption was overstated. Indeed, comparison of AB isotherms onto raw extracted



(a) $20 \text{ mg.L}^{-1} < [AB]_0 < 200 \text{ mg.L}^{-1}$; (b) $200 \text{ mg.L}^{-1} < [AB]_0 < 1000 \text{ mg.L}^{-1}$

Figure 12: Spectral scanning of residual solutions after adsorption of AB onto the extracted cellulose saturated with BB dye.

cellulose and saturated one with BB dye by taking into account of the cationic dye release, shows quantitatively that adsorption of anionic dye has not enhanced as much as if each dye adsorbed on its own site. However, the shape of the isotherms appears slightly modified because the saturated surface state is different from the raw one.

A previous study [24], which investigated competitive adsorption of anionic and cationic dyes onto multiwalled carbon nanotubes showed that, in binary dye system, synergistic effect was observed. So cationic dye presented favorite adsorption on the sorbent due to hydrophobic effect and π - π bonds and authors demonstrate that the surface adsorbed cationic dye has a positive charge, which would serve as the sites for electrostatic attraction to negatively charged of the anionic dye. This result is concluded from macroscopic adsorption study, which shows an improvement in the adsorption of anionic dye in binary system compared to single one. However it should be noted that the author did not take into account the dyes mutual interaction in mixture, which leads to drawn wrong conclusions for the adsorption process.

Conclusion

The extracted cellulose from *P. oceanica*, locally available and low cost material, presented as good alternative adsorbents for Basic blue 9 (BB) and Acid blue 25 (AB) removal from aqueous solutions. Equilibrium data agreed well with Langmuir isotherm model with adsorption capacity of 0.955 and 0.370 mmol.g⁻¹ for BB and AB, respectively. Extracted cellulose before and after dye adsorption was characterized by FTIR spectroscopy and XPS analysis indicating that adsorption of both dyes could be achieved mainly on external fibre surface with different adsorption interactions. It seems that electrostatic and hydrogen bonding involving respectively charged groups and OH groups takes place in the interaction between dye species and the fibre surface.

On the other hand, the adsorption of AB dye onto extracted cellulose previously saturated by the BB was investigated and showed that the determination of residual AB concentrations cannot be carried directly from the maximum wavelength of the dye, due to its interaction in solution with BB. Indeed, the formation of molecular complex detected at wavelengths 510 and 705 nm when the anionic (AB) and the cationic (BB) dyes were taken in equimolar proportions, has been revealed and has made it more difficult to determine the binary mixture composition after adsorption. The AB adsorption isotherm in this case was investigated and showed that the presence of BB on the surface affects the shape of the adsorption isotherm of the anionic dye but not the adsorbed amounts.

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