

What Is Acid Rain and How Is It formed?

Seyedtaghi Mirmohammadi*

Department of Occupational Health Faculty of Health, Mazandaran University of Medical Sciences, Iran

Editorial

Acid rain is a result of air pollution. When any type of energy is burnt, lots of different chemicals are produced. The bank that comes from a fire or the smothers that come out of an auto exhaust do not just contain the sooty slate patches that you can see-they also contains lots of unnoticeable feasts that can be indeed more dangerous to our terrain [1].

Power stations, manufactories and buses all burn energies and thus they all produce contaminating feasts. Some of these feasts (especially nitrogen oxides and sulphur dioxide) reply with the bits driblets of water in shadows to form sulphuric and nitric acids. The rain from these shadows also falls as veritably weak acid-which is why it's known as "acid rain". The release of sulphur dioxide can also do naturally when a flash point erupts [2].

Acid rain is one of the consequences of air pollution. It occurs when emigrations from manufactories, buses or hitting boilers contact with the water in the atmosphere. These emigrations contain nitrogen oxides, sulphur dioxide and sulphur trioxide, which when mixed with water come sulphurous acid, nitric acid and sulfuric acid. This process also occurs naturally through stormy eruptions [3].

Acid rain describes any form of precipitation that contains high levels of nitric and sulfuric acids. It can also occur in the form of snow, fog, and tiny bits of dry material that settle to Earth. Normal rain is slightly acidic, with a pH of 5.6, while acid rain generally has a pH between 4.2 and 4.4 [4].

The performing acids are rained to earth as rain or snow with veritably negative consequences on the one hand the damage to nature in the form of acidification of soils, lakes and swell with consequent damage to terrestrial and marine foliage and fauna. On the other hand, acid rain also causes erosion of metallic rudiments- structures, islands, halls and other structures-and the destruction of mortal heritage made of limestone- structures and major structures, statues, puppets [5].

The only way to stop acid rain is to reduce emigrations that beget it. This involves laying on renewable energy sources and reducing the use of fossil energies in the artificial and automotive sector and in the diurnal life of every citizen. Pure water is considered a neutral substance. It has a pH of 7. Clean, or unpolluted, rain is naturally slightly acidic with a pH of 5.6. Acid rain is any precipitation that has a pH lower than 5.6. Typically, the pH of acid rain is between 2 and 4.5.

Have you ever put vinegar (aqueous acetic acid) on your french fries? If so, you consumed an acid (vinegar) with a pH of about 2 - 2.5. Fries with vinegar are pretty delicious. So if acid rain and vinegar both have low pH, then acid rain can't be all that bad, right? Wrong! Acid rain can disturb forest growth, kill fish and aquatic plants, and harm buildings, too [6].

Acid rain occurs substantially in the Northern Hemisphere-- the more bucolic, dirtier half of the globe. Winds can sweep up emigrations from high smokestacks and carry adulterants far from their original sources, crossing state lines and public borders in the process. Acid rain may not have the complete global range of hothouse feasts, but it's a

transboundary, and thus transnational, issue [7].

Acid rain, also known as acid deposit, is caused by emigrations of sulphur dioxide (SO₂) and nitrogen oxides (NO_x) from power shops, buses and manufactories. Natural sources like tinderboxes, timber fires and lightning strikes also add to the man- made pollution. SO₂ and NO_x come acids when they enter the atmosphere and reply with water vapour [8]. The performing sulfuric and nitric acids can fall as wet or dry grounds. Wet deposit is rush acid rain, snow, sleet or fog. Dry deposit falls as acidic particulates or feasts. Timbers calculate on their soil's buffering capacity to cover them from acid rain. Acidic waters draw out soil poisons like aluminium. Trees take in the toxic substances, and runoff dumps it in lakes, gutters and aqueducts [9]. Acid rain also dissolves helpful minerals and nutrients like calcium, magnesium and potassium before trees can absorb them. Acid rain infrequently kills a timber outright but rather stunts its growth through times of soil declination. Nutrient privation and exposure to poisons make trees more likely to trip in storms or die in cold rainfall [10].

Conflict of Interest

None

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None

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*Corresponding author: Seyedtaghi Mirmohammadi, Department of Occupational Health Faculty of Health, Mazandaran University of Medical Sciences, Iran, E-mail: Seyedtaghi.Mirmohammadi@gmail.com

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